**Gas Laws Quiz– General**

1. Why is it reasonable for us to assume that gas molecules don’t experience intermolecular forces? (5 pt)
2. What is an ideal gas, and how is it related to real gases? (5 pt)
3. If I move a 2.0 L balloon from the surface of a pool (pressure = 1.00 atm) to 1 meter under the water (pressure = 1.03 atm), what will its new volume be? (5 pt)
4. A burp usually has a volume of about 0.020 L. Assuming the body temperature of a human is 37.5 degrees Celsius and the pressure inside the stomach is about 1.015 atm, how many moles of gas are in a typical burp? (5 pt)
5. Why does the volume of a gas go up as its temperature increases? (5 pt)